

Synthèse : S

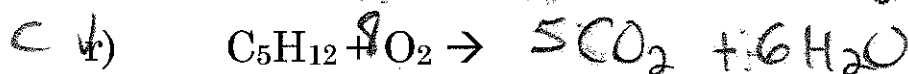
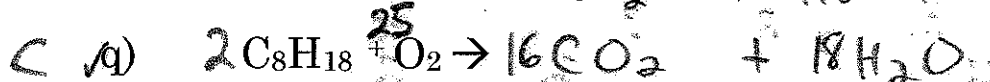
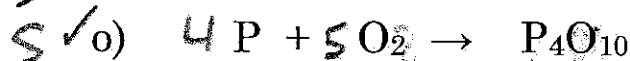
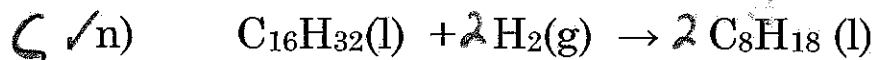
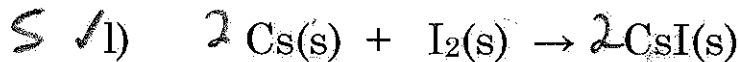
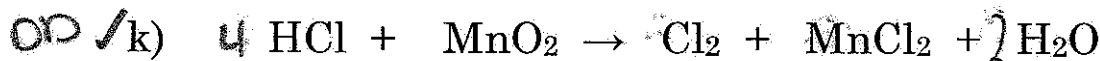
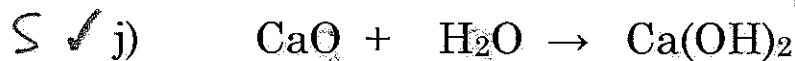
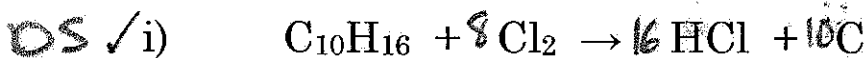
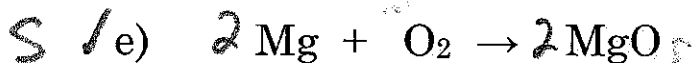
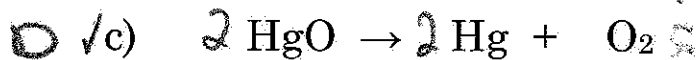
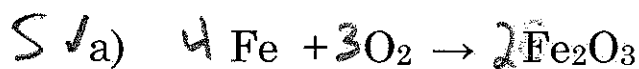
Combustion : C

décomposition : D déplacement simple : DS

Déplacement
double :

DD

1. Identifiez et balancez les équations suivantes :



18
50

2. Équilibrez les équations suivantes :

- ✓ 1. $2 \text{Al} + 6 \text{HCl} \rightarrow 2 \text{AlCl}_3 + 3 \text{H}_2$
- ✓ 2. $\text{Al} + 3 \text{NaOH} \rightarrow 3 \text{Na} + \text{Al(OH)}_3$
- ✓ 3. $2 \text{Al(OH)}_3 \rightarrow \text{Al}_2\text{O}_3 + 3 \text{H}_2\text{O}$
- ✓ 4. $2 \text{NaOH} + \text{Al}_2\text{O}_3 \rightarrow 2 \text{NaAlO}_2 + \text{H}_2\text{O}$
- ✓ 5. $2 \text{NaAlO}_2 + 2 \text{H}_2\text{O} \rightarrow 2 \text{NaOH} + \text{Al(OH)}_3$
- ✓ 6. $3 \text{FeCl}_2 + 2 \text{K}_3\text{Fe(CN)}_6 \rightarrow 2 \text{Fe}_3(\text{Fe(CN)}_6)_2 + 6 \text{KCl}$
- ✓ 7. $\text{C}_6\text{H}_{12}\text{O}_6 \rightarrow 2 \text{C}_2\text{H}_5\text{OH} + 2 \text{CO}_2$
- ✓ 8. $\text{C}_2\text{H}_5\text{OH} + 3 \text{O}_2 \rightarrow 2 \text{CO}_2 + 3 \text{H}_2\text{O}$
- ✓ 9. $2 \text{CH}_3\text{OH} + 3 \text{O}_2 \rightarrow 2 \text{CO}_2 + 4 \text{H}_2\text{O}$
- ✓ 10. $\text{CaC}_2 + 2 \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{C}_2\text{H}_2$
- ✓ 11. $\text{CH}_4 + 2 \text{O}_2 \rightarrow \text{CO}_2 + 2 \text{H}_2\text{O}$
- ✓ 12. $2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$
- ✓ 13. $3 \text{H}_2 + \text{N}_2 \rightarrow 2 \text{NH}_3$
- ✓ 14. $(\text{NH}_4)_2\text{SO}_4 + \text{Ca(OH)}_2 \rightarrow \text{CaSO}_4 + 2 \text{NH}_3 + 2 \text{H}_2\text{O}$
- ✓ 15. $2 \text{ZnS} + 3 \text{O}_2 \rightarrow 2 \text{ZnO} + 2 \text{SO}_2$
- ✓ 16. $\text{Cr}_2\text{O}_3 + 2 \text{Al} \rightarrow 2 \text{Cr} + \text{Al}_2\text{O}_3$
- ✓ 17. $2 \text{C}_{17}\text{H}_{35}\text{COONa} + \text{Ca} \rightarrow 2 \text{Na} + (\text{C}_{17}\text{H}_{35}\text{COO})_2\text{Ca}$
- ✓ 18. $\text{Al(OH)}_3 + 3 \text{HCl} \rightarrow \text{AlCl}_3 + 3 \text{H}_2\text{O}$
- ✓ 19. $5 \text{CO} + \text{I}_2\text{O}_5 \rightarrow 5 \text{CO}_2 + \text{I}_2$
- ✓ 20. $\text{Ni} + 4 \text{CO} \rightarrow \text{Ni(CO)}_4$
- ✓ 21. $2 \text{C}_8\text{H}_{18} + 17 \text{O}_2 \rightarrow 8 \text{CO} + 18 \text{H}_2\text{O} + 8 \text{CO}_2$
- ✓ 22. $2 \text{NH}_3 + \text{CO}_2 \rightarrow (\text{NH}_2)_2\text{CO} + \text{H}_2\text{O}$
- ✓ 23. $\text{Fe}_2\text{O}_3 + 3 \text{CO} \rightarrow 2 \text{Fe} + 3 \text{CO}_2$
- ✓ 24. $\text{Cu} + 4 \text{HNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2 \text{NO}_2 + 2 \text{H}_2\text{O}$
- ✓ 25. $4 \text{NH}_3 + 5 \text{O}_2 \rightarrow 4 \text{NO} + 6 \text{H}_2\text{O}$

18
8
8
34